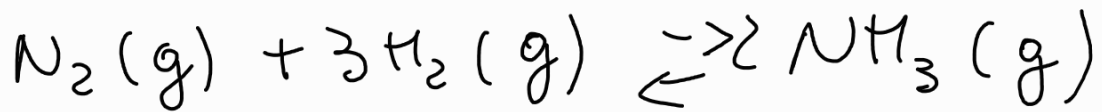


1

ESOTERMICO QUINDI POSSO DIRE
CHE IL CALORE È UN PRODOTTO
 $\Delta H < 0$



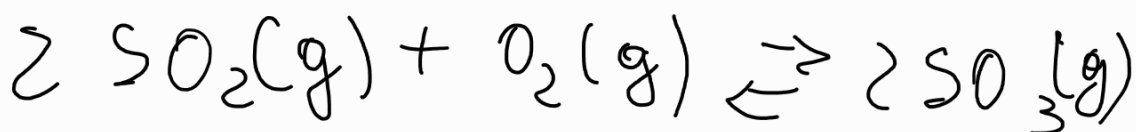
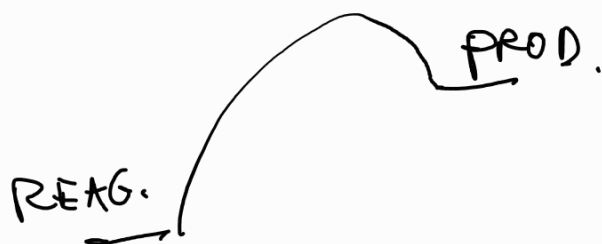
4 MOL DI GAS 2 MOL DI GAS

- A) AUMENTO VOLUME FAVORISCE I REAGENTI
(NO)
- B) AUMENTO PRESSIONE N_2 FAVORISCE
I PRODOTTI (SI)
- C) AUMENTO PRESSIONE DI NH_3
FAVORISCE REAGENTI (NO)
- D) AUMENTO TEMPERATURA FAVORISCE
REAGENTI (NO)

B

②

ENDOTERMICA $\Delta H > 0$
CALORE E' REAGENTE

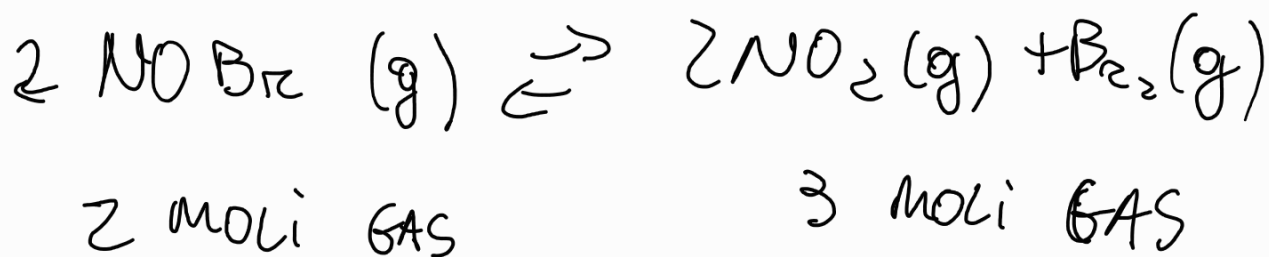


3 MOLİ GAS

2 MOLİ
GAS

- A) AUMENTO VOLUME SPOSTA VERSO
REAGENTI (NO)
- B) DIMIN. PRESSIONE SO_2 SPOSTA
VERSO REAGENTI (NO)
- C) AUMENTO PRESSIONE TOTALE
SPOSTA VERSO PRODOTTI (MENO GAS)
(SI) $\boxed{1/C}$
- D) DIMIN. TEMPERATURA SPOSTA
A REAGENTI (NO)

3



ENDOTERMICA CALORE REAGENTE

- a) AGGIUNTA NOBr SPOSTA VERSO PRODOTTI (SI)
- b) AGGIUNTA Br₂ SPOSTA VERSO REAGENTI (NO)
- c) DIM. TEMP. SPOSTA VERSO REAGENTI (NO)
- d) AUMENTO VOLUME SPOSTA VERSO PRODOTTI (PIU' MOLI GAS) (SI)

A

4



1.2

3.6

0

0

-X

-3X

+X

+X

1.2 - X

3.6 - 3X

X

X

$$3.6 - 3X = 0.90 \text{ M}$$

$$X = 0.90$$

$$K_c = \frac{[\text{S}_2\text{Cl}_2][\text{CCl}_4]}{[\text{CS}_2][\text{Cl}_2]^3}$$

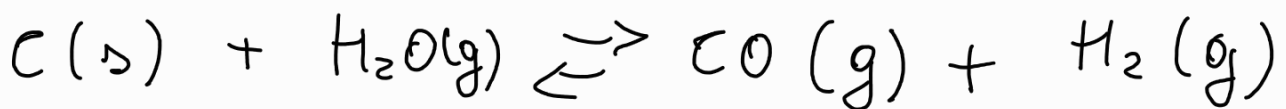
$$= \frac{0.9 \cdot 0.9}{0.3 \cdot 0.9^3} = 3.70$$

17

5

$$P_{H_2O} = \frac{36 \cdot R \cdot T}{V} = 15.9 \text{ atm}$$

$$K_p = \frac{P_{CO} \cdot P_{H_2}}{P_{H_2O}}$$



-	15.9	0	0
---	------	---	---

-	-X	+X	+X
---	----	----	----

-	15.9 - X	X	X
---	----------	---	---

$$K_p = 2.85 = \frac{X^2}{15.9 - X}$$

$$X^2 + 2.85 \cdot X - 45.31 = 0$$

$$X = 5.45 = P_{H_2} = P_{CO}$$

A

⑥

K PICCOLA DUNQUE EQUILIBRIO
SPOSTATO VERSO REAGENTI

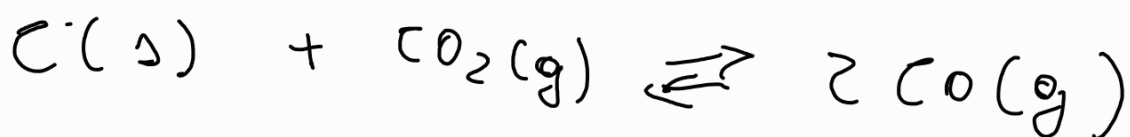
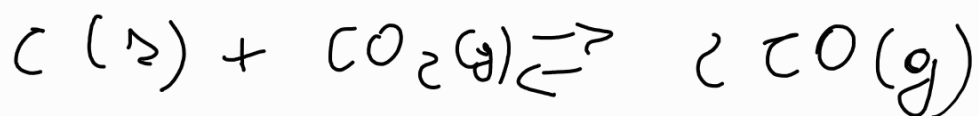
$$R \rightleftharpoons P \quad K = \frac{[P]}{[R]} \Rightarrow [P] = K [R]$$

$$[P] = 10^{-4} [R]$$

✓

(7)

$$P_{CO_2} = 0.5 \text{ atm} \quad P_{CO} = 0.34 \text{ atm}$$



$$\sim \quad 0.5 \quad 0$$

$$\sim \quad -X \quad +2X$$

$$\sim \quad 0.5 - X \quad 2X$$

$$2X = 0.34 \text{ atm}$$

$$X = 0.17 \text{ atm}$$

Quindi :

$$K_p = \frac{P_{CO}^2}{P_{CO_2}} = \frac{0.34^2}{(0.5 - 0.17)} = 0.35$$

A

